

Fax to

November 5, 2010

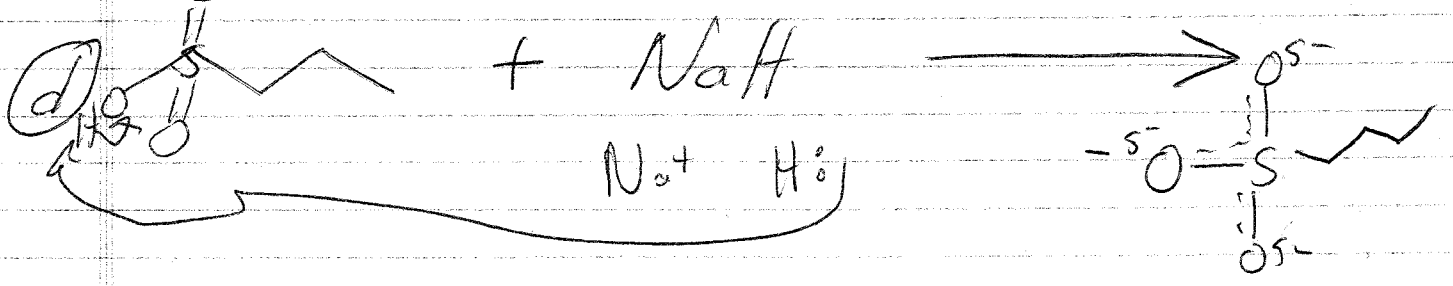
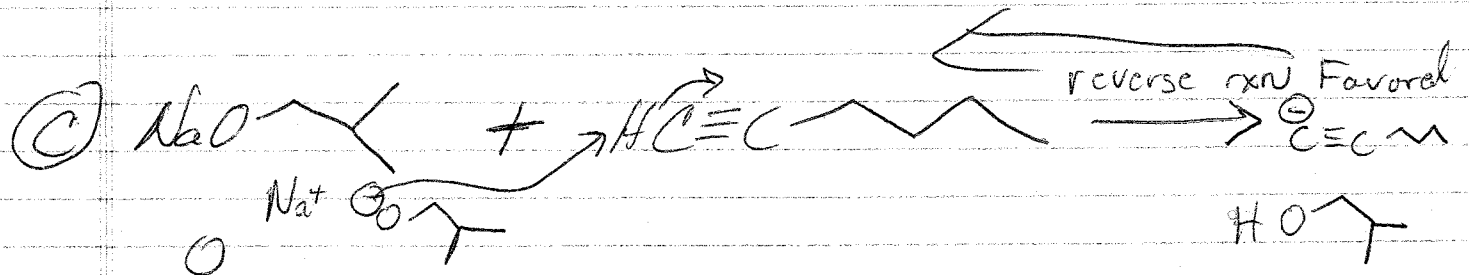
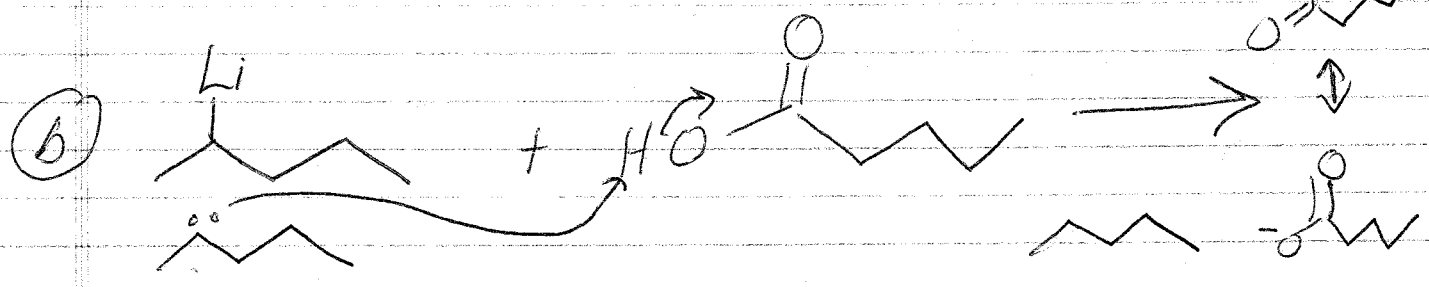
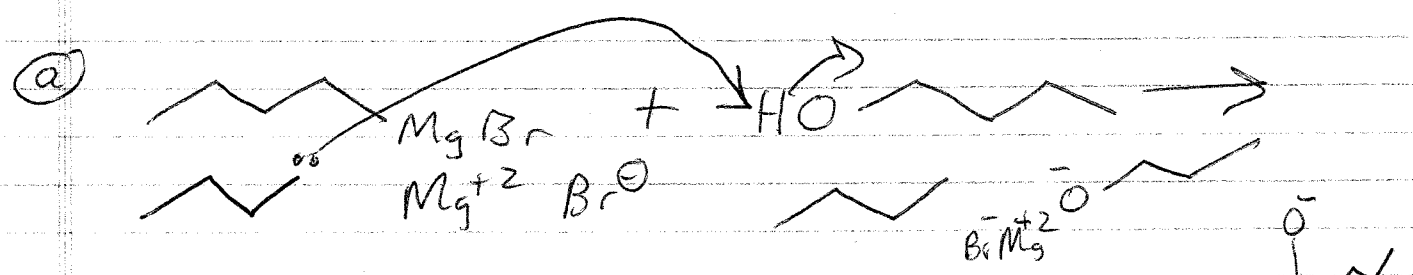
1/13

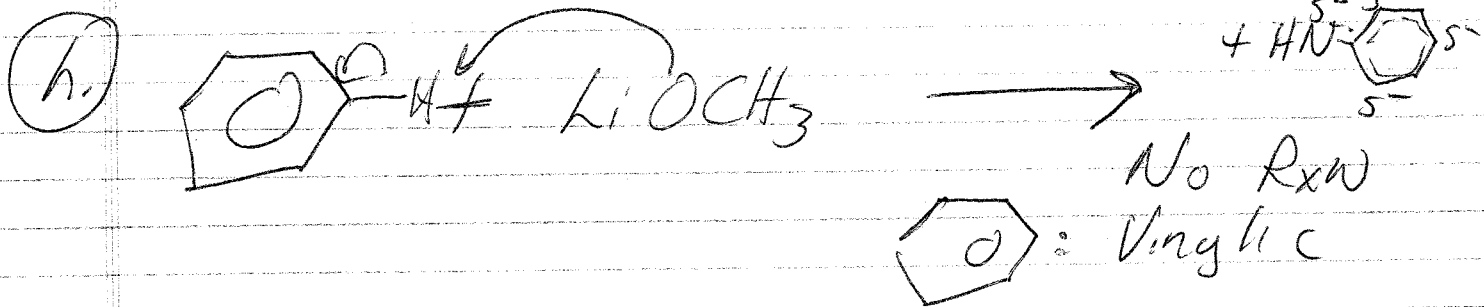
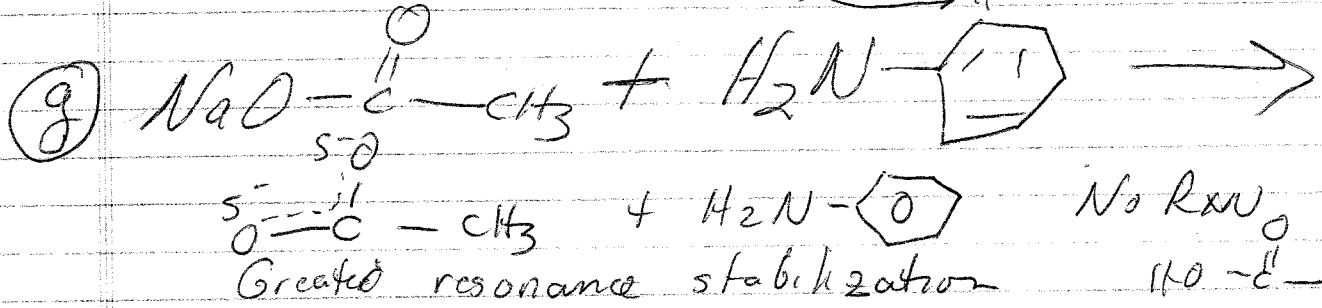
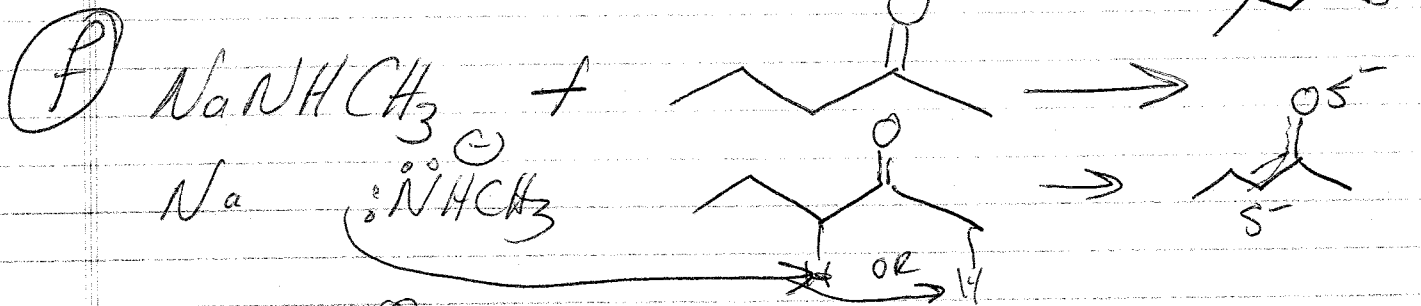
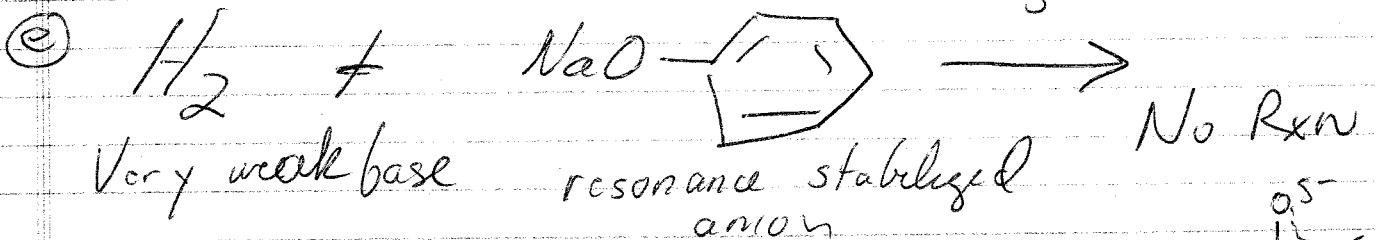
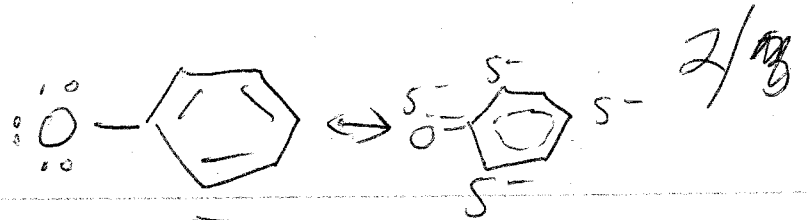
M. Nerz

Chem 211

Solutions will be posted 5 bonus pts for printing & turning in first page of solution.

1) Write mechanism (arrow formalism) leading to major products for each of the following reactions. If you do not believe a reaction will occur, simply write no reaction.

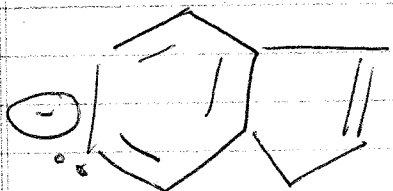




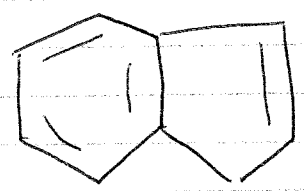
(2) For any of the above structures where resonance of an anion is relevant, write resonance forms

See Above

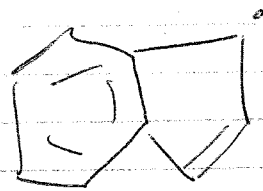
3.



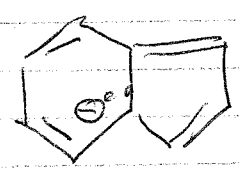
vinylic  
 No resonance  
 Strong base  
 conjugate acid w



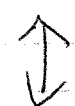
allylic



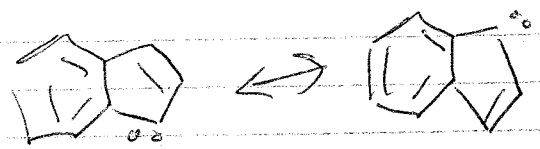
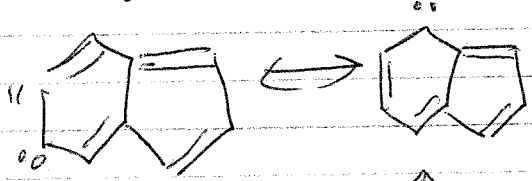
High  
 or  
 extensive  
 resonance  
 stabilizer



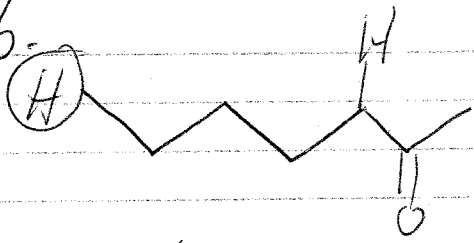
Conjugate  
 Acid



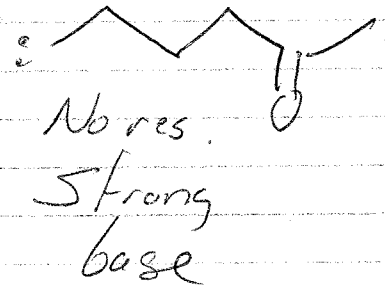
"stronger"



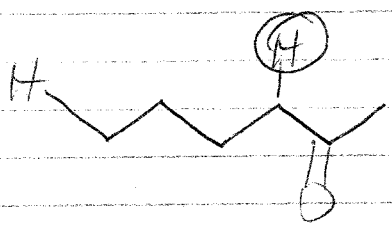
3. b.



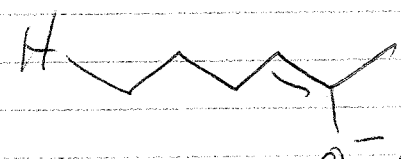
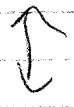
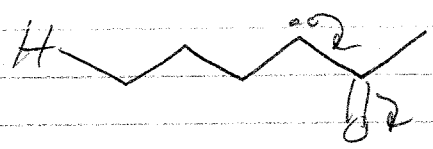
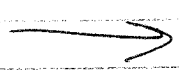
weak acid



No res.  
Strong base

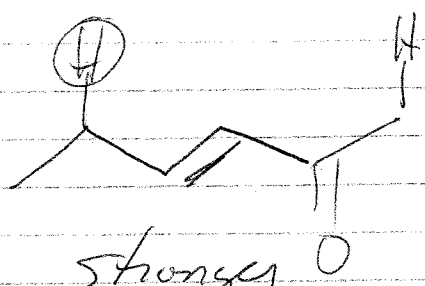


Stronger acid

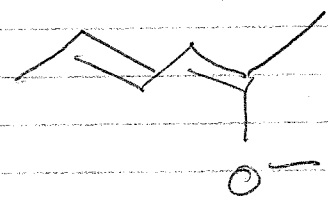
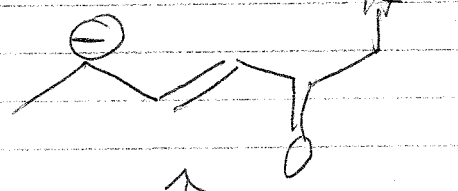


Resonance stabilizes charge on oxygen  
weaker base

c

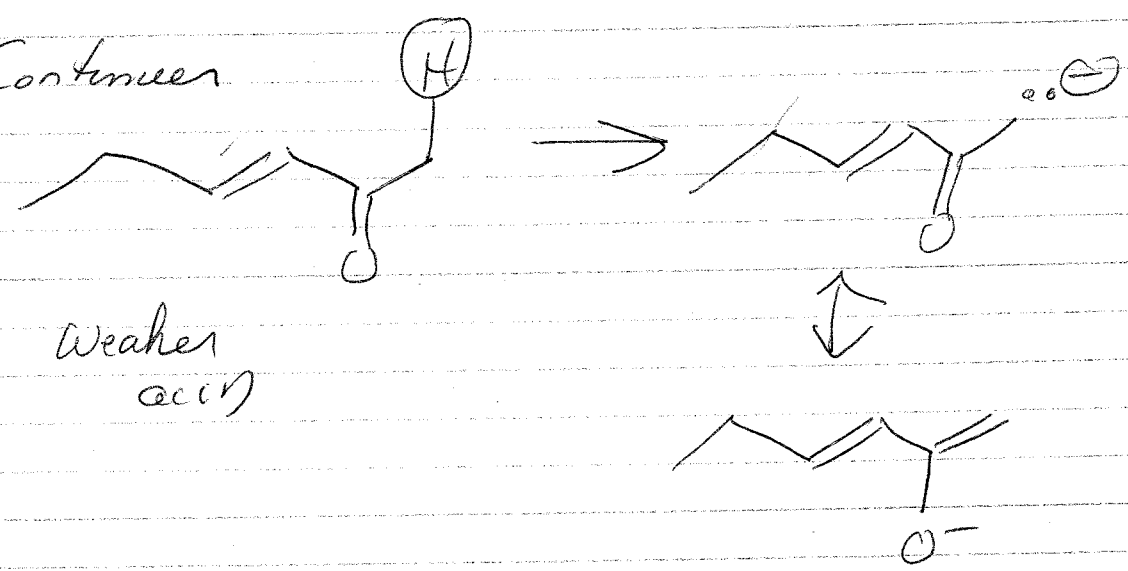


Stronger acid



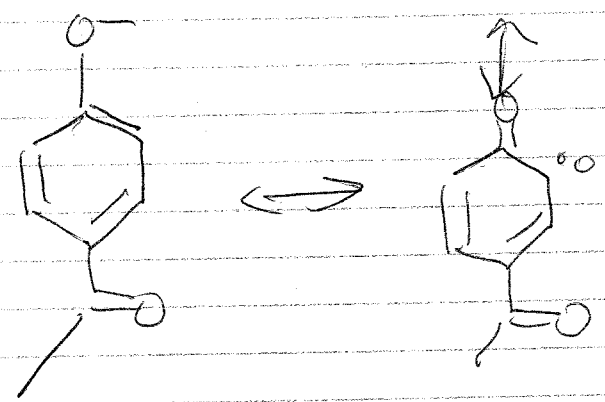
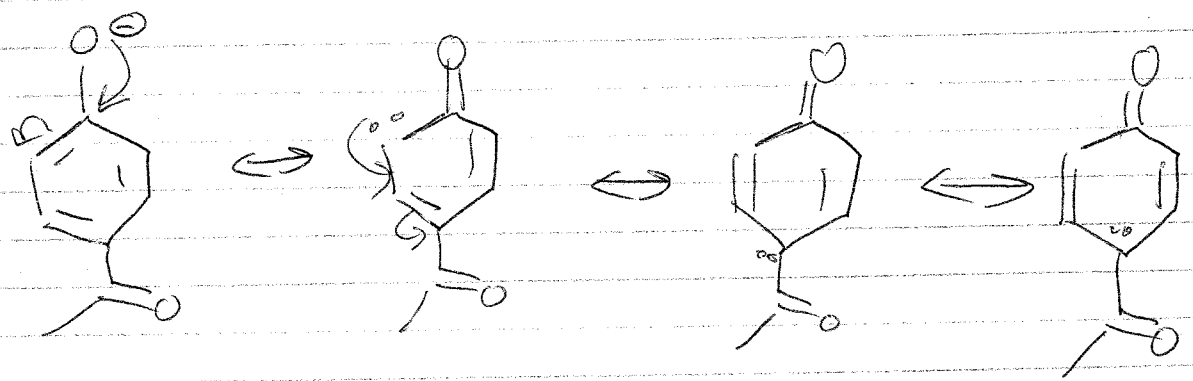
Greater resonance stabilizer  
weaker base

C. Continue

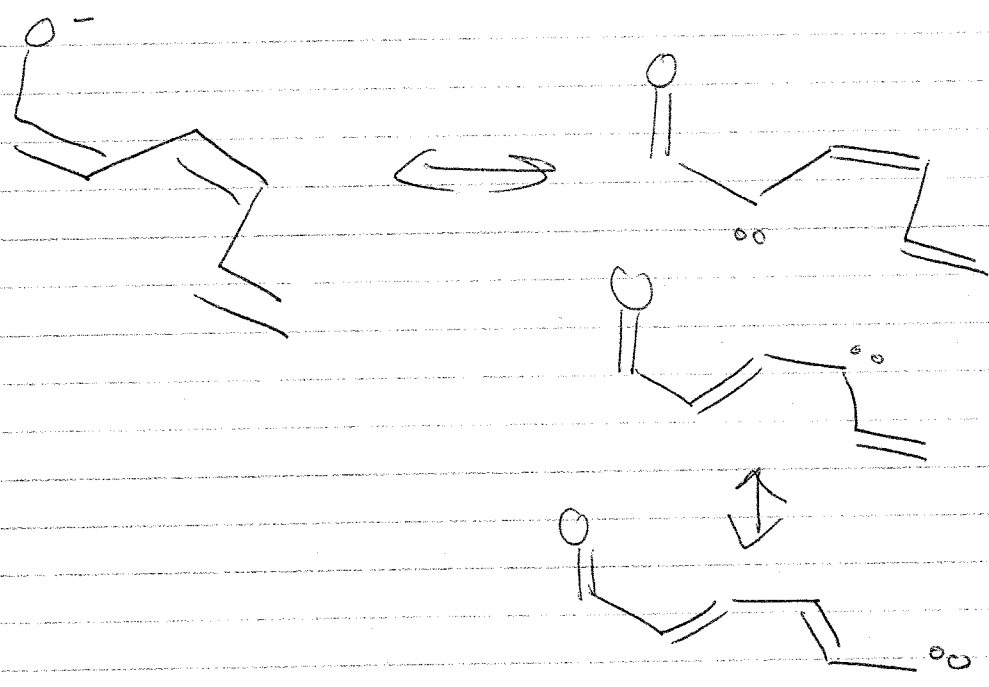


less resonance stabilization  
stronger base

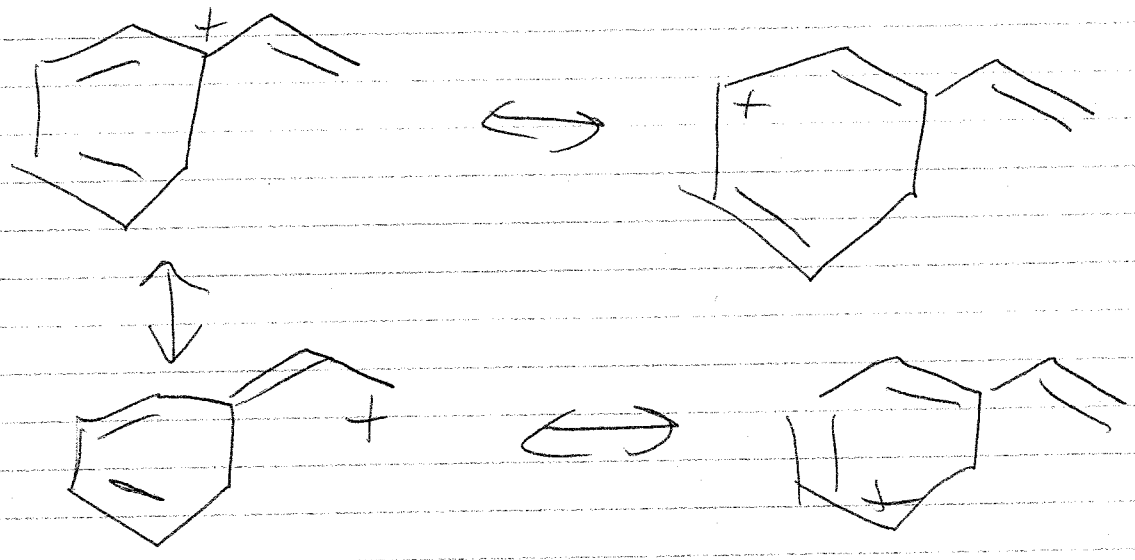
4.



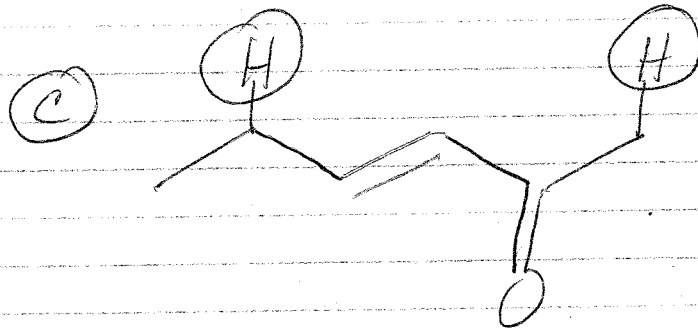
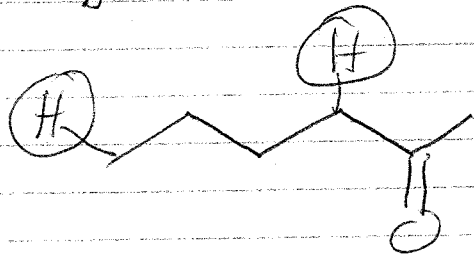
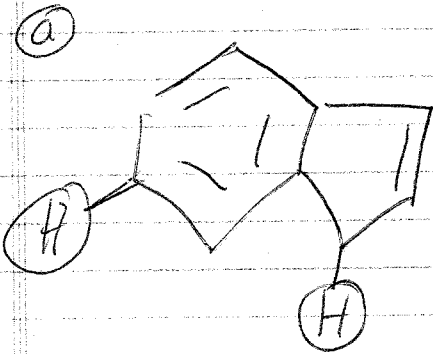
b.



c.



3. Which of the following protons is more acidic. Write resonance forms if relevant in your answer.



4. Write all resonance forms for the following species.

